

GOVERNMENT DEGREE COLLEGE (AUTONOMOUS), SIDDIPET

SYLLABUS W.E.F THE ACADEMIC YEAR 2017-18

B.Sc. I YEAR, SEMESTER – I

SUBJECT: CHEMISTRY

PAPER-I

**Unit-I (Inorganic Chemistry) 15h (1 hr/week)**

**S1-I-1. s-block elements: 2 h**

General Characteristics of groups I and II elements, Diagonal relationship between Li and Mg, Be and Al

**S1-I-2. p-block elements 1: 7 h**

Group-13: Synthesis and structure of diborane and higher Boranes ( $B_4H_{10}$  and  $B_5H_9$ ),

Boron nitrogen compounds ( $B_3N_3H_6$  and BN), Lewis acid nature of  $BX_3$

Group – 14: Carbides-Classification – ionic, covalent, interstitial – synthesis. Structures and reactivity. Industrial application. Silicones – Preparation – a) direct silicon process b) use of Grignard reagent c) aromatic silylation. Classification – straight chain, cyclic and cross-linked.

Group – 15: Nitrides – Classification – ionic, covalent and interstitial. Reactivity – hydrolysis. Preparation and reactions of hydrazine, hydroxyl amine, phosphazenes.

**S1-I-3. General Principles of Inorganic qualitative analysis 6 h**

Anion analysis: Theory of sodium carbonate extract, classification and reactions of Anions

$CO_3^{2-}$ , Cl<sup>-</sup>, Br<sup>-</sup>,  $SO_4^{2-}$ ,  $PO_4^{3-}$ ,  $BO_3^{3-}$ ,  $CH_3COO^-$ ,  $NO_3^-$ .

Cation Analysis: Principles involved - Solubility product, common ion effect, general discussion for the separation and identification of group I individual cations ( $Hg_2^+$ , Ag<sup>+</sup>, Pb<sup>+</sup>) with flow chart and chemical equations. Principle involved in separation of group II & IV cations.

General discussion for the separation and identification of group II ( $Hg_2^+$ , Pb<sup>2+</sup>, Bi<sup>3+</sup>, Cd<sup>2+</sup>, Sb<sup>2+</sup>), III (Al<sup>3+</sup>, Fe<sup>3+</sup>), IV ( $Mn^{2+}$ , Zn<sup>2+</sup>) individual cations with flow chart and chemical equations. Application of concept of hydrolysis in group V cation analysis.

General discussion for the separation and identification of group V individual cations

(Ba<sup>2+</sup>, Sr<sup>2+</sup>, Ca<sup>2+</sup>) with flow chart and chemical equations. Theory of flame test.

Identification of Group VI cations ( $Mg^{2+}$ ,  $NH_4^+$ ).

**Unit - II (Organic Chemistry) 15h (1 hr/week)**

**S1-O-1: Structural Theory in Organic Chemistry 6 h**

**Bond polarization:** Factors influencing the polarization of covalent bonds, electro negativity – inductive effect. Application of inductive effect (a) Basicity of amines (b) Acidity of carboxylic acids (c) Stability of carbonium ions. Resonance - Mesomeric effect, application to (a) acidity of phenol. (b) acidity of carboxylic acids and basicity of anilines. Stability of carbo cations, carbanions and free radicals. Hyper conjugation and its application to stability of carbonium ions, Free radicals and alkenes.

**Types of organic reactions:** Addition reactions- electrophilic, nucleophilic and free radical. Substitution reactions – electrophilic, nucleophilic and free radical. Elimination and Rearrangement reactions– Examples.

### **S1-O-2: Acyclic Hydrocarbons**

**6 h**

**Alkanes**– Methods of preparation: Corey-House reaction, Wurtz reaction, from Grignard reagent, Kolbe synthesis. Chemical reactivity - inert nature, free radical substitution, Halogenation example- reactivity, selectivity and orientation.

**Alkenes** - Preparation of alkenes (with mechanism) (a) by dehydration of alcohols (b) dehydrohalogenation of alkyl halides (c) by dehalogenation of 1,2 dihalides, Zaitsev's rule. Properties: Addition of Hydrogen – heat of hydrogenation and stability of alkenes. trans-addition of halogen and its mechanism. Addition of HX, Markonikov's rule, addition of H<sub>2</sub>O, HOX, H<sub>2</sub>SO<sub>4</sub> with mechanism and addition of HBr in the presence of peroxide (anti – Markonikov's addition). Oxidation (cis – additions) – hydroxylation by KMnO<sub>4</sub>, OsO<sub>4</sub>, trans addition- peracids (via epoxidation), hydroboration, ozonolysis – location of double bond. Dienes – Types of dienes, reactions of conjugated dienes – 1,2 and 1,4 addition of HBr to 1,3 – butadiene and Diels – Alder reaction.

**Alkynes**– Preparation by dehydrohalogenation of vicinal dihalides, dehalogenation of tetrahalides. Physical Properties: Acidity of terminal alkynes (formation of metal acetylides) preparation of higher alkynes, Chemical reactivity – electrophilic addition of X<sub>2</sub>, HX, H<sub>2</sub>O (tautomerism), Oxidation (formation of enediol, 1,2 diones and carboxylic acids) and reduction (Metal-ammonia reduction, catalytic hydrogenation)

### **S1-O-3: Alicyclic Hydrocarbons**

**3 h**

Nomenclature, preparation by Freund's method, Dickmann, heating dicarboxylic metal salts. Properties – reactivity of cyclo propane and cyclo butane by comparing with alkanes. Stability of cycloalkanes – Baeyer strain theory, Sachse and Mohr predictions and Pitzer strain theory. Conformational structures of cyclopentane, cyclohexane.

### **Unit-III (Physical Chemistry)**

**15 h (1 hr/week)**

#### **S1-P-1: Atomic structure and elementary quantum mechanics**

**6 h**

Black body radiation, heat capacities of solids, Rayleigh Jeans law, Planck's radiation law, photoelectric effect, Limitations of classical mechanics, Compton effect, De Broglie's hypothesis. Heisenberg's uncertainty principle, Schrodinger's wave equation and its importance. Physical interpretation of the wave function, significance of  $\psi$  and  $\psi^2$ , a particle in a box, energy levels, wave functions and probability densities. Schrodinger wave equation for H-atom. Separation of variables, radial and angular functions (only equation), hydrogen like wave functions, quantum numbers and their importance.

#### **S1-P-2: Gaseous State**

**5 h**

Deviation of real gases from ideal behavior. van der Waals equation of state. Critical phenomenon. PV isotherms of real gases, continuity of state. Andrew's isotherms of CO<sub>2</sub>. The van der Waal's equation and critical state. Derivation of relationship between critical constants and van der Waal's constants. The law of corresponding states, reduced equation of states. Joule Thomson effect and inversion temperature of a gas. Liquefaction of gases: i) Linde's method based on Joule Thomson effect ii) Claude's method based on adiabatic expansion of a gas.

#### **S1-P-3: Liquid State**

**4 h**

Intermolecular forces, structure of liquids (qualitative description). Structural differences between solids, liquids and gases. Surface tension and its determination using stalagmometer. Viscosity of a liquid and determination of coefficient of viscosity using Ostwald viscometer. Effect of temperature on surface tension and coefficient of viscosity of a liquid (qualitative treatment only). Liquid crystals, the mesomorphic state:

Classification of liquid crystals into Smectic and Nematic, differences between liquid crystal and solid / liquid. Application of liquid crystals as LCD devices.

**Unit – IV (General Chemistry)**

**15 h (1 hr/week)**

**S1-G-1 Chemical Bonding**

**11 h**

Ionic solids- lattice and solvation energy, solubility of ionic solids, Fajan's rule, polarity and polarizability of ions, covalent nature of ionic bond, covalent bond - Common hybridization and shapes of molecules.

**Molecular orbital theory:** Shapes and sign convention of atomic orbitals. Modes of overlapping. Concept of  $\sigma$  and  $\pi$  bonds. Criteria for orbital overlap. LCAO concept.

Types of molecular orbitals- bonding, antibonding and non bonding. MOED of homonuclear diatomics -  $H_2$ ,  $N_2$ ,  $O_2$ ,  $O_2^-$ ,  $O_2^{2-}$ ,  $F_2$  (unhybridized diagrams only) and heteronuclear diatomics CO, CN-NO,  $NO^+$  and HF. Bond order, stability and magnetic properties.

**S1-G-2 Evaluation of analytical data**

**4 h**

Significant figures, accuracy and precision. Errors-classification of errors- determinate and indeterminate errors, absolute and relative errors, propagation of errors in mathematical operations – addition, subtraction, division and multiplication (with respect to determinate errors).

**References:**

**Unit- I**

1. Principles of Inorganic Chemistry by Puri, Sharma and Kalia Vishal Publications 1996.
2. Concise Inorganic Chemistry by J.D. Lee 3<sup>rd</sup> edn.
3. Basic Inorganic Chemistry by F.A.Cotton, G.Wilkinson and Paul.L. Gaus 3<sup>rd</sup> edn Wiley Publishers 2001. Chem.
4. Vogel's Qualitative Inorganic Analysis by Svehla
5. Inorganic Chemistry Principles of structure and reactivity by James E.Huhey, E.A. Keiter and R.L. Keiter 4<sup>th</sup> edn.
6. Chemistry of the elements by N.N.Greenwood and A. Earnshaw Pergamon Press 1989.
7. Inorganic Chemistry by Shriver and Atkins 3<sup>rd</sup> edn Oxford Press 1999.
8. Qualitative analysis by Welcher and Hahn.
9. Textbook of Inorganic Chemistry by R Gopalan
10. College Practical chemistry by V K Ahluwalia, Sunitha Dhingra and Adarsh Gulati

**Unit- II**

1. Text book of organic chemistry by Morrison and Boyd.
2. Text book of organic chemistry by Graham Solomons.
3. Text book of organic chemistry by Bruice Yuranis Powla.
4. Text book of organic chemistry by Soni.
5. General Organic chemistry by Sachin Kumar Ghosh.
6. Text book of organic chemistry by C N pillai

**Unit III**

1. Principles of physical chemistry by Prutton and Marron.
2. Text Book of Physical Chemistry by Soni and Dharmahara..
3. Text Book of Physical Chemistry by Puri and Sharma.

4. Text Book of Physical Chemistry by K. L. Kapoor.
5. Physical Chemistry through problems by S.K. Dogra.
6. Text Book of Physical Chemistry by R.P. Verma.
7. Elements of Physical Chemistry by Lewis Glasstone.

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1. Principles of Inorganic Chemistry by Puri, Sharma and Kalia Vishal Publications 1996.
2. Concise Inorganic Chemistry by J.D. Lee 3<sup>rd</sup> edn.
3. Basic Inorganic Chemistry by F.A.Cotton, G.Wilkinson and Paul.L. Gaus 3<sup>rd</sup> edn Wiley Publishers 2001. Chem
4. Analytical chemistry by G. L. David Krupadanam, D. Vijaya Prasad, K. Varaprasada Rao, K.L.N. Reddy and C. Sudhakar

**Laboratory Course**

**45h (3 h / week)**

**Paper I Qualitative Analysis - I**

***I. Preparations:***

1. Tetrammine copper (II) sulphate,
2. Potash alum  $KAl(SO_4)_2 \cdot 12H_2O$ ,
3. Bis (dimethylglyoximate) nickel(II)

***II. Analysis of two anions (one simple and one interfering)***

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**SYLLABUS W.E.F THE ACADEMIC YEAR 2017-18**

**B.Sc. I YEAR, SEMESTER – II**

**SUBJECT: CHEMISTRY**

**PAPER-II**

**Unit-I (Inorganic Chemistry)**

**S2-I-1 p-block Elements -II**

**15 h (1 hr/week)**

**7 h**

**Oxides:** Types of oxides (a) Normal- acidic, basic amphoteric and neutral (b) Mixed(c) sub oxide d) peroxide e) superoxide. Structure of oxides of C, N, P, S and Cl - reactivity, thermal stability, hydrolysis.

**Oxy acids:** Structure and acidic nature of oxyacids of B, C, N, P, S and Cl. Redox properties of oxyacids of Nitrogen: HNO<sub>2</sub> (reaction with FeSO<sub>4</sub>, KMnO<sub>4</sub>, K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>), HNO<sub>3</sub> (reaction with H<sub>2</sub>S, Cu), HNO<sub>4</sub> (reaction with KBr, Aniline), H<sub>2</sub>N<sub>2</sub>O<sub>2</sub> (reaction with KMnO<sub>4</sub>). Redox properties of oxyacids of Potassium: H<sub>3</sub>PO<sub>2</sub> (reaction with HgCl<sub>2</sub>), H<sub>3</sub>PO<sub>3</sub> (reaction with AgNO<sub>3</sub>, CuSO<sub>4</sub>).

Redox properties of oxyacids of Sulphur: H<sub>2</sub>SO<sub>3</sub> (reaction with KMnO<sub>4</sub>, K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>), H<sub>2</sub>SO<sub>4</sub> (reaction with Zn, Fe, Cu), H<sub>2</sub>S<sub>2</sub>O<sub>3</sub> (reaction with Cu, Au), H<sub>2</sub>SO<sub>5</sub> (reaction with KI, FeSO<sub>4</sub>), H<sub>2</sub>S<sub>2</sub>O<sub>8</sub> (reaction with FeSO<sub>4</sub>, KI)

**Interhalogens-** classification- general preparation- structures of AB, AB<sub>3</sub>, AB<sub>5</sub> and AB<sub>7</sub> type and reactivity. Poly halides- definition and structure of ICl<sub>2</sub>-, ICl<sub>4</sub>- and I<sub>3</sub>-. Comparison of Pseudohalogens with halogens.

### **S2-I-2 Chemistry of Zero group elements**

**2 h**

General preparation, structure, bonding and reactivity of Xenon compounds – Oxides, Halides and Oxy-halides. Clathrate compounds and Anomalous behavior of He (II)

### **S2-I-3 Chemistry of d-block elements**

**6 h**

Characteristics of d-block elements with special reference to electronic configuration variable valence, ability to form complexes, magnetic properties & catalytic properties. Stability of various oxidation states and SRP Comparative treatment of second and third transition series with their 3d analogues. Study of Ti, Cr and Cu triads. Titanium triad – electronic configuration and reactivity of +3 and +4 states – oxides and halides. Chromium triad – reactivity of +3 and +6 states. Copper triad – reactivity of +1, +2 and +3 states.

## **Unit – II (Organic chemistry)**

**15 h (1 hr/week)**

### **S2-O-1: Aromatic Hydrocarbons**

**7h**

Concept of aromaticity – definition, Huckel's rule – application to Benzenoids and Non – Benzenoids (cyclopropenyl cation, cyclopentadienyl anion and tropylium cation). Preparations: From acetylene, phenols, benzene carboxylic acids and sulphonic acids. Reactions - General mechanism of electrophilic substitution, mechanism of nitration, sulphonation, and halogenation, Friedel Craft's alkylation (polyalkylation) and acylation. Orientation of aromatic substitution - Definition of ortho, para, and meta directing groups. Ring activating and deactivating groups with examples. Orientation – (i) activating groups: Amino, methoxy and alkyl groups. (ii) Deactivating groups - carboxy, nitro, nitrile, carbonyl and sulphonic acid & halo groups.

### **S2-O-2: Arenes and Polynuclear Aromatic Hydrocarbons**

**3 h**

Preparation of alkyl benzenes by Friedel Craft's alkylation, Friedel Craft's acylation followed by reduction, Wurtz-Fittig reaction. Chemical reactivity: Ring substitution reactions, side chain substitution reactions and oxidation.

Polynuclear hydrocarbons – Structure of naphthalene and anthracene (Molecular Orbital diagram and resonance energy) Reactivity towards electrophilic substitution. Nitration and sulphonation as examples.

### **S2-O-3: Halogen compounds**

**5 hrs**

Nomenclature and classification: alkyl (primary, secondary, tertiary), aryl, aralkyl, allyl, vinyl, benzyl. Chemical reactivity - reduction, formation of RMgX, Nucleophilic substitution reactions – classification into  $S_N1$  and  $S_N2$ . Mechanism and energy profile diagrams of  $S_N1$  and  $S_N2$  reactions. Stereochemistry of  $S_N2$  (Walden Inversion) 2-bromobutane,  $S_N1$  (Racemisation) 1-bromo-1-phenylpropane explanation of both by taking the example of optically active alkyl halide. Structure and reactivity – Ease hydrolysis - comparison of alkyl, vinyl, allyl, aryl, and benzyl halides.

**Unit – III (Physical Chemistry)**

**15 h (1 hr/week)**

**S2-P-1: Solutions**

**5 h**

Liquid - liquid mixtures, ideal liquid mixtures, Raoult's and Henry's laws. Non ideal systems. Azeotropes HCl-H<sub>2</sub>O and C<sub>2</sub>H<sub>5</sub>OH - H<sub>2</sub>O systems. Fractional distillation, Partially miscible liquids- Phenol – Water, Trimethyl amine – Water and Nicotine – Water systems. Lower upper consolute temperatures. Effect of impurity on consolute temperature. Immiscible liquids and steam distillation. Nernst distribution law. Calculation of the partition coefficient. Applications of distribution law with solvent extraction.

**S2-P-2: Dilute Solutions & Colligative Properties**

**5 h**

Dilute Solutions, Colligative Properties, Raoult's law, relative lowering of vapour pressure, molecular weight determination. Osmosis - laws of osmotic pressure, its measurement, determination of molecular weight from osmotic pressure. Elevation of boiling point and depression of freezing point. Derivation of relation between molecular weight and elevation in boiling point and depression in freezing point. Experimental methods for determining various colligative properties. Abnormal molar mass, Van't Hoff factor, degree of dissociation and association of solutes.

**S2-P-3: Solid state Chemistry**

**5 h**

Laws of Crystallography – (i) Law of Constancy of interfacial angles (ii) Law of Symmetry, Symmetry elements in crystals (iii) Law of rationality of indices. Definition of space lattice, unit cell. Bravais Lattices and Seven Crystal systems (a brief review). X-ray diffraction by crystals; Derivation of Bragg's equation, Determination of structure of NaCl, KCl & CsCl (Bragg's method and Powder method).

**Unit – IV (General Chemistry)**

**15 h (1 hr/week)**

**S2-G-1: Theory of Quantitative Analysis**

**5 h**

**Volumetric Analysis:** Introduction, standard solutions, indicators, end point, titration curves, Types of titrations: i) neutralization titration- principle, theory of acid base indicators, titration curves and selection of indicators- strong acid - strong base, strong acid –weak base, weak acid- strong base and weak acid –weak base.

Gravimetric analysis- Introduction, nucleation, precipitation, growth of precipitate, filtration and washing, drying and incineration of precipitate, coprecipitation and post precipitation. Determination of Ni<sup>2+</sup>

**S3-G-2: Theories of bonding in metals:**

**5 h**

Valence bond theory, Explanation of metallic properties and its limitations, Free electron theory, thermal and electrical conductivity of metals, limitations, Band theory, formation of bands, explanation of conductors, semiconductors n-type and p-type, extrinsic & intrinsic semiconductors, and insulators.

**S2-G-3: Material Science**

**5 h**

Classification of materials- classification as metals, ceramics, organic polymers, composites, biological materials etc. The property of super conductivity of materials.

Super conducting materials- elements, alloys and compounds. Properties of super conductors- zero resistivity, Meisener effect and thermal properties. Composites meaning of composites, advanced composites, classification –particle reinforced fiber reinforced and structural composites general characters of composite materials-Particle reinforced composites – large particle and dispersion- strengthened composite. Fiber reinforced composites (continuous and discontinuous fiber composites).

## References

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1. Principles of Inorganic Chemistry by Puri, Sharma and Kalia Vishal Publications 1996.
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5. Physical Chemistry through problems by S.K. Dogra.
6. Elements of Physical Chemistry by Lewis and Glasstone.
7. Material science by Kakani & Kakani

### Unit IV

1. Vogel's Text Book of Quantitative Analysis by G.H.Jeffery, J.Bassett, J.Mendham and R.C. Denney 5<sup>th</sup> edn Addison Wesley Longman Inc. 1999.
2. Quantitative Analysis by Day and Underwood Prentice Hall (India) VI Edn..
3. Nano: The Essentials by T. Pradeep, McGraw-Hill Education.
4. Chemistry of nanomaterials: Synthesis, Properties and applications by CNR Rao et.al.
5. Nanostructured Materials and Nanotechnology, edited by Hari Singh Nalwa, Academic Press
6. College Practical chemistry by V K Ahluwalia, Sunitha Dhingra and Adarsh Gulati

## Laboratory Course 45hrs (3 h / week)

### Paper II - Qualitative Analysis - II

#### I Semi micro analysis of mixtures

Analysis of two anions and two cations in the given mixture.

Anions:  $\text{CO}_3^{2-}$

$\text{SO}_3^{2-}$ ,  $\text{S}^{2-}$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{CH}_3\text{COO}^-$ ,  $\text{NO}_3^-$ ,  $\text{PO}_4^{3-}$ ,  $\text{BO}_3^{3-}$ ,  $\text{SO}_4^{2-}$

Cations:  $\text{Ag}^+$ ,  $\text{Pb}^{2+}$ ,  $\text{Hg}^+$ ,  $\text{Hg}_2^{2+}$

$\text{Pb}^{2+}$ ,  $\text{Bi}^{3+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{As}^{3+/5+}$ ,  $\text{Sb}^{3+/5+}$ ,  $\text{Sn}^{2+/4+}$

$\text{Al}^{3+}$ ,  $\text{Cr}^{3+}$ ,  $\text{Fe}^{3+}$

$\text{Zn}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$

$\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$

$\text{Mg}^{2+}$ ,  $\text{NH}_4^+$